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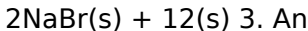
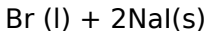
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Note: Not all of the  
reactions will occur. For  
those that do not, write

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no reaction. 1. A piece of copper is dropped into a container of water. No reaction 2.

Liquid bromine is added to a container of sodium iodide crystals.



3. An aluminum strip is immersed in a solution of silver nitrate.

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Free elements, no  
matter how complex  
the molecule, have an  
oxidation number  
(valence or charge)  
equal to zero. Oj. Fe.

T~ He < /1 1 ~ ~ L-td.

8. The following are  
diatomic or polyatomic  
elements in nature  
which must be  
committed to memory.

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28, Practice-Pg. 49 •

Formation of a Gas -

Pg. 50, Practice-Pg. 51

• Acid Base

Neutralization

Reactions- Pg. 53,

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acids and bases do not

break apart and are in

net ionic equation. •

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SnCO tin(II) carbonate  
sodium hydrogen  
carbonate NaHCO 1. 2.  
3. 4. 5. 6. 8. 9. 10. 11.  
vanadium(V) oxide H<sub>2</sub>O  
dihydrogen monoxide

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(NH) CO ammonium  
oxalate 4 224  
polonium(VI)  
thiocyanate Po(SCN)6  
tetrphosphorus  
decaoxicle P O 4 10 12.  
13. 15. 16. 17. 18. 19.  
manganese(VII) oxide  
copper(II) dihydrogen  
phosphate

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Answers Chapter

3  
Correlating questions  
(ii) from: Rodney  
Schmitz (2007) Moore  
County Schools, NC  
Reaction Prediction – 5

For each of the  
following eight  
reactions, in part (i)  
write a BALANCED  
equation and in part (ii)  
answer the question  
about the reaction.

**KEY to 5 Reaction**

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Determine the  
oxidation number of  
the underlined  
element:  $\text{KMnO}_4$  Since  
K is an alkali metal, its  
charge must be  $1+$   
Oxygen is  $2-$  but there  
are four of them,  
therefore, 4 times  $2-$   
equals  $8-$  If  $1+$  and  
 $8-$  are added Ultimate  
Chemical Equation

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K is an alkali metal, its  
charge must be  $1+$   
Oxygen is  $2-$  but there  
are four of them,  
therefore, 4 times  $2-$

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equals 8— If 1+ and  
8— are added The  
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